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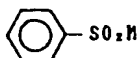
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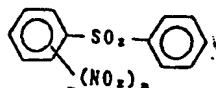
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(64) A method of producing alkali metal benzenesulfonates.

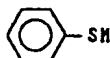
(67) A method of producing alkali metal benzenesulfonates having the general formula of



wherein M represents Na or K, which comprises: reacting nitrophenylphenyl sulfones having the general formula of



wherein n is an integer of 1 or 2, with alkali metal thiophenolates having the general formula of



in solvents.

The reaction produces also nitrophenylphenyl sulfides as by-products in equimolar amounts to the alkali metal benzenesulfonates, and the oxidation of the sulfides provides the starting nitrophenylphenyl sulfones.

A method of producing the nitrophenylphenyl sulfones is also provided wherein the sulfides are oxidized with hydrogen peroxide in the presence both of water-soluble tungstates or molybdates and of phase transfer catalysts in a two-phase heterogeneous solvent.

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A Method of Producing Alkali Metal Benzenesulfates

10 This invention relates to a method of producing alkali metal benzenesulfates, and also to a method of producing nitrophenylphenyl sulfones usable as starting materials in the production of alkali metal benzenesulfates.

15 Alkali metal benzenesulfates are useful intermediate raw materials for the production of organic sulfones and sofoxides important in organic synthetic industry.

20 Up to date a variety of methods are already known to produce alkali metal benzenesulfates. These prior methods include, for example, a method in which benzenesulfonyl chloride is reduced with sodium sulfite as described in A. D. Barnard et al., J. Chem. Soc., 1957, 4673. However, according to this method, sodium sulfinate is produced in solutions together with water-soluble compounds such as sodium sulfate, sodium chloride or sodium benzenesulfonate, so that
25 it is very difficult to separate sodium benzenesulfinate

from the by-products by crystallization on account of closeness of the solubilities in water of all these compounds. Therefore, it is generally preferred that the resultant solution is then acidified with a mineral acid to convert the alkali metal benzenesulfonates into the free acid, and then the free acid is separated by filtration or is extracted with organic solvents. This method is, however, at disadvantage from the industrial standpoint in that it handles the free acid which is chemically rather unstable as well as it needs the aforesaid additional steps that incur higher production cost.

A further method is also known in which benzene is sulfonated with sulfur dioxide in the presence of aluminum chloride as a catalyst, as described in Nippon Kagaku Zasshi, 1968, 89(8), 810. This method needs the catalyst in amounts equivalent to benzene used, and the catalyst cannot be recovered after the reaction, which makes the method uneconomical. A still further method is known in which 2,4-dinitrophenylphenyl sulfone is reacted with potassium hydroxide to provide potassium benzenesulfonate, as described in C. N. Kharash et al., J. Org. Chem., 19, 1704 (1954). However, this method has disadvantages in that the separation of potassium benzenesulfonate from by-products, potassium 2,4-dinitrophenolate is difficult.

As described above, the methods hitherto known have

disadvantages in many respects, and there is known no method that is advantageously applicable to the industrial production of alkali metal benzenesulfonates.

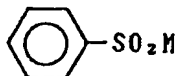
The present inventors have made extensive investigations to establish an improved method of producing alkali metal benzenesulfonates more efficiently and less expensively than the prior methods, and found out that alkali metal benzenesulfonates are readily obtained in high yields by the reaction of nitrophenylphenyl sulfones with alkali metal thiophenolates in solvents.

It is, therefore, an object of the invention to provide a novel method of producing alkali metal benzenesulfonates.

It is also an object of the invention to provide a novel method of producing nitrophenylphenyl sulfones which are usable as starting materials in the production of alkali metal benzenesulfonates.

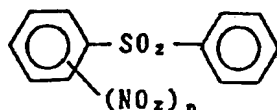
According to the invention, there is provided a method of producing alkali metal benzenesulfonate having the general formula of

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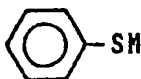


wherein M represents Na or K, which comprises: reacting nitrophenylphenyl sulfones having the general formula of

25



wherein n is an integer of 1 or 2, with alkali metal thiophenolates having the general formula of



5 in solvents.

The nitrophenylphenyl sulfones used in the invention include 2- and 4-nitrophenylphenyl sulfone and 2,4-dinitrophenylphenyl sulfone. The alkali metal thiophenolates used in the invention include sodium and potassium salts, which
10 provide sodium and potassium benzenesulfinate, respectively, in almost quantitatively by the reaction with the aforesaid nitrophenylphenyl sulfones. In the reaction, the alkali metal thiophenolates are usually used in amounts of about 0.8-1.2 moles per, preferably in amounts equivalent to, the
15 nitrophenylphenyl sulfones used.

The reaction is carried out in solvents usually composed of a mixture of water and organic solvents that dissolve therein nitrophenylphenyl sulfones used and are
20 inert to the reaction. The organic solvents usable may be either water-soluble or water-insoluble, and include, for example, lower aliphatic alcohols of 1-5 carbons such as methanol, ethanol or isopropanol, aliphatic hydrocarbons such as hexane or heptane, aromatic hydrocarbons that may
25 carry one or more of lower alkyls of each 1-3 carbons such as benzene, toluene, xylene, ethylbenzene or isopropyl-

benzene, lower alkyl esters of lower aliphatic carboxylic acids having in total 3-8 carbons such as methyl acetate, ethyl acetate or methyl propionate, nitrated aromatic hydrocarbons such as nitrobenzene or nitrotoluene.

5 The reaction may be carried out either in homogeneous solvents or in two-phase heterogeneous solvents depending upon the organic solvents used. By way of example, a mixture of water and water-soluble organic solvents such as methanol or ethanol is a homogeneous solvent, whereas a mixture of
10 water and water-insoluble or slightly water-soluble organic solvents such as ethyl acetate, toluene, nitrobenzene or nitrotoluene is a two-phase heterogeneous solvent.

Water-soluble or water-miscible aliphatic alcohols such as methanol or ethanol may be used alone as a solvent
15 in the reaction since such alcohols are capable of dissolving not only thiophenol and alkali metal hydroxides but also nitrophenylphenyl sulfones therein. However, thiophenol and alkali metal hydroxides are reacted with each other therein to produce in situ alkali thiophenolates, and the reaction
20 produces also water, so that if the alcohols alone are used as a solvent, they inevitably include water therein.

The reaction may be carried out in manners that are not specifically limited, but it is generally preferred that the reaction is carried out by adding a solution of alkali
25 thiophenolates in water, methanol or hydrous methanol,

to a solution of nitrophenylphenyl sulfones in organic solvents as hereinbefore mentioned under stirring, and then the resultant mixture is stirred, if desired, at elevated temperatures, for about 1-10 hours. The reaction temperature is usually in the range of about 30-100°C. When the reaction temperature is too low, the reaction proceeds too slowly to be industrially workable, whereas when the reaction temperature is too high, undesired side reactions take place to decrease the yields of the objective compounds and the selectivity of the reaction. The alkali metal thiophenolates may be prepared by, for example, reacting thiophenol with sodium or potassium hydroxide in equimolar amounts in water, methanol or a mixture of these.

When a two-phase heterogeneous solvent is used, the reaction is preferably carried out in the presence of phase transfer catalysts, which are already known per se. The phase transfer catalysts usable in the invention include, for example, tetraalkylammoniums such as tetramethylammonium chloride, tetraethylammonium chloride, tetrabutylammonium bromide or trioctylmethylammomium chloride, tetraalkylammonium hydrogen sulfates such as tetrabutylammonium hydrogen sulfate, benzyltrialkylammoniums such as benzyltrimethylammonium chloride, benzyltriethylammonium chloride, benzyldimethylaurylammonium chloride or benzyldimethyltetradecylammonium chloride, and dibenzylalkylammoniums

such as dibenzyltrimethylammonium chloride or dibenzyl-diethylammonium chloride. The use of the phase transfer catalysts accelerates the reaction and increases the yields of the alkali metal benzenesulfonates.

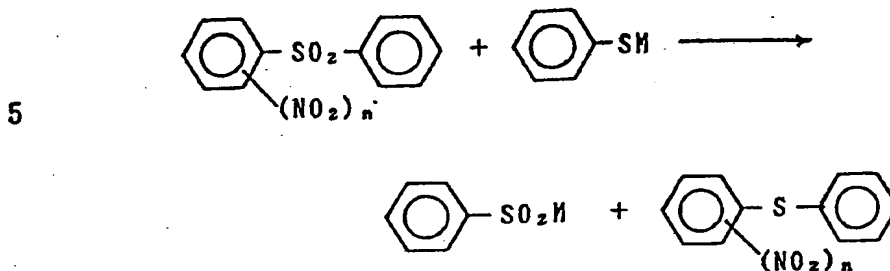
5 The phase transfer catalysts are used in amounts of about 0.1-5 %, preferably 0.2-1 % by weight, based on the weight of the alkali metal thiophenolates used. The use of too small amounts is ineffective to accelerate the reaction, while the use of too large amounts is uneconomical.

10 When the reaction is carried out in a homogeneous solvent, the solvent is concentrated after the reaction, water and organic solvents such as toluene are added to the resultant concentrate, and the aqueous solution is separated and concentrated to dryness, to provide alkali metal benzene-
15 sulfonates as white crystals. As by-products, nitrophenyl-phenyl sulfides are recovered from the organic solution.

 When the reaction is carried out in a heterogeneous solvent, the reaction mixture is separated into an aqueous solution and an organic solution after the reaction, the
20 aqueous solution is washed with organic solvents, for example, toluene, and is then concentrated to dryness, to provide alkali metal benzenesulfonates as white crystals. The nitrophenylphenyl sulfides are recovered from the organic solution.

25 The reaction of nitrophenylphenyl sulfones with alkali

metal thiophenolates according to the invention is shown below.



wherein M represents Na or K, and n is an integer of 1 or 2.

10 As shown above, the reaction produces, in addition to the alkali metal benzenesulfates, nitrophenylphenyl sulfides, as by-products, which can be oxidized with oxidizing agents to the corresponding nitrophenylphenyl sulfones initially used. In this regard, nitrophenylphenyl sulfones

15 that have one or two nitro groups on one of the aromatic nuclei provide both alkali metal benzenesulfates and nitrophenylphenyl sulfides in high yields, in almost equimolar amounts.

20 A variety of oxidizing agents are usable for oxidizing nitrophenylphenyl sulfides into the corresponding nitrophenylphenyl sulfones. The oxidizing agents usable in the invention include, for example, hydrogen peroxide, peracids such as peracetic acid, hydroperoxides, halogens such as chlorine or bromine, ozone, oxygen with transition metal catalysts,

25 potassium peroxysulfate, potassium permanganate, dinitrogen

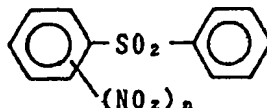
tetroxide, sodium metaperiodate, osmium (VIII) oxide, ruthenium (VIII) oxide, sodium or potassium dichromate and nitric acid.

5 However, hydrogen peroxide is most preferred as oxidizing agents since the oxidation of the nitrophenylphenyl sulfides therewith provides the corresponding nitrophenylphenyl sulfones almost quantitatively.

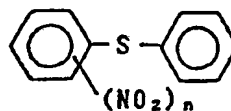
10 Therefore, as an important aspect of the invention, the recovered nitrophenylphenyl sulfides produced as by-products can be reused in the reaction after the oxidation. More specifically, the recovered nitrophenylphenyl sulfides are oxidized to the nitrophenylphenyl sulfones, and the sulfones are anew reacted with the alkali metal thiophenolates, to provide the alkali metal benzenesulfinates
15 and again the nitrophenylphenyl sulfides. Therefore, the nitrophenylphenyl sulfones as raw materials are needed to initially carry out the reaction, however, if needed, only supplementary amounts of nitrophenylphenyl sulfones are needed in the succeeding reactions, and hence according to
20 the invention, the alkali metal benzenesulfinates can be produced much less expensively than the prior methods ever known.

25 According to the invention, there is further provided a novel method of producing nitrophenylphenyl sulfones by the oxidation of nitrophenylphenyl sulfides.

According to the invention, there is further provided a method of producing nitrophenylphenyl sulfones having the general formula of



wherein n is an integer of 1 or 2, which comprises: oxidizing nitrophenylphenyl sulfides having the general formula of



wherein n is an integer of 1 or 2, with hydrogen peroxide in the presence both of oxidizing catalysts selected from the group consisting of water-soluble tungstates and molybdates and of phase transfer catalysts in a two-phase heterogeneous solvent.

15

The oxidation reaction is carried out in a two-phase heterogeneous solvent, which is composed of water and water-insoluble or slightly water-soluble organic solvents. Any water-insoluble or slightly water-soluble organic solvents are usable when they are capable of dissolving therein nitrophenylphenyl sulfides and are inert to the reaction. The organic solvents usable include, for example, aliphatic hydrocarbons such as hexane, heptane or octane, aromatic hydrocarbons such as benzene, toluene, xylene, ethylbenzene or isopropylbenzene, alicyclic hydrocarbons

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such as cyclohexane or ethylcyclohexane, and nitrated aromatic hydrocarbons such as nitrobenzene or nitrotoluene. Alkyl esters of acetic acid or higher carboxylic acids are also usable as organic solvents, such as ethyl acetate, propyl acetate, butyl acetate, amyl acetate or ethyl propionate.

In the oxidation reaction, hydrogen peroxide is used as oxidizing agents in amounts of not less than about 2 moles, preferably in the range of about 2-2.5 moles, per mole of nitrophenylphenyl sulfides used. There is no need to use hydrogen peroxide in large excess amounts.

The tungstates and molybdates used as oxidizing catalysts in the reaction are water-soluble, and include alkali metal salts such as lithium, sodium or potassium salts, alkaline earth metal salts such as magnesium salts, and ammoniums. Preferred catalysts are therefore lithium tungstate, sodium tungstate, potassium tungstate, magnesium tungstate, ammonium tungstate, sodium molybdate, potassium molybdate and ammonium molybdate.

The oxidizing catalyst is used in amounts of about 0.5-5 % by weight, preferably of about 1-2 % by weight, based on the weight of the nitrophenylphenyl sulfides used. The use of too small amounts is ineffective to accelerate the reaction, whereas the use of too large amounts is uneconomical.

In the oxidizing reaction of the invention, phase transfer catalysts are used together with the oxidizing catalysts. The phase transfer catalysts hereinbefore described are also usable in the reaction. Preferred
5 phase transfer catalysts used include, for example, tetraalkylammoniums such as tetramethylammonium chloride, tetraethylammonium chloride, tetrabutylammonium bromide or trioctylmethylammonium chloride, tetraalkylammonium hydrogen sulfates such as tetrabutylammonium hydrogen
10 sulfate, benzyltrialkylammoniums such as benzyltrimethylammonium chloride, benzyltriethylammonium chloride, benzyl-dimethyl-laurylammonium chloride or benzyldimethyltetradecylammonium chloride, and dibenzylalkylammoniums such as
15 dibenzyl-dimethylammonium chloride or dibenzyl-diethylammonium chloride .

The phase transfer catalysts are used usually in amounts of about 0.1-3 % by weight, preferably of about 0.2-1 % by weight, based on the weight of nitrophenylphenyl sulfides used. The use of too small amounts is ineffective to
20 accelerate the reaction, whereas the use of too large amounts is uneconomical.

The oxidation reaction is preferably carried out in acidic or neutral conditions by adding a small amount of inorganic strong acids such as sulfuric acid or hydrochloric
25 acid to the reaction mixture. Namely, the reaction is

carried out preferably at a pH of not more than about 7, and most preferably at a pH of about 1-7. However, when ammonium hydrogen sulfates such as tetrabutylammonium hydrogen sulfate are used as a phase transfer catalyst, there is no need to add an inorganic strong acid to the reaction mixture to adjust the pH thereof at acidic regions since the ammonium hydrogen sulfates themselves are strongly acidic.

Although the invention is not limited in theory or mechanism of the reaction, the oxidizing catalysts are converted into acid forms, i.e., tungstic acid or molybdic acid in acidic or neutral conditions, and the acids react with the phase transfer catalysts and transfer from the water phase to the organic phase in the reaction mixture, thereby to effectively catalyze the oxidation of the nitrophenylphenyl sulfides into the corresponding nitrophenylphenyl sulfones. More specifically, it is likely that the tungstates or molybdates are oxidized into peroxotungstic acid or peroxomolybdic acid, respectively, in acidic or neutral conditions, and the peroxyacids react with the phase transfer catalysts to transfer to the organic phase, whereby to effectively catalyze the oxidation of the sulfides to the sulfones.

On the other hand, when the reaction is carried out in alkaline conditions, the oxidizing catalysts remain in

the water phase in the form of the salts in the reaction mixture, and it is likely that the salts do not react with the phase transfer catalysts, thereby being prevented from transferring from the water phase to the organic phase where the substantial oxidation reaction is carried out.

The oxidizing reaction is carried out usually at temperatures of about 50-90°C, preferably at about 60-80°C. When the reaction temperature is too low, the reaction proceeds too slowly, and when the reaction temperature is too high, the selectivity of the reaction and the yields of the sulfones are low.

The reaction manners are not specifically limited, however, it is preferred that the reaction is carried out by dissolving nitrophenylphenyl sulfides in the organic solvents and adding to the resultant solution the oxidizing catalysts, the phase transfer catalysts, and then, if necessary, an inorganic strong acid, and then by adding dropwise thereto aqueous solution of hydrogen peroxide under stirring at temperatures of about 60-90°C. After the completion of the reaction, the reaction mixture is separated into an organic solution and aqueous solution, and the organic solution is concentrated to dryness, to provide nitrophenylphenyl sulfones in high yields. The resultant sulfones may be purified by recrystallization.

The invention will be understood more readily with

reference to the following examples; however these examples are intended to illustrate the invention only and are not to be construed as limitation to the invention.

5

EXAMPLE 1

(i) Production of Sodium Benzenesulfinate and Recovery of 2-Nitrophenylphenyl Sulfide

An amount of 26.3 g (0.1 mole) of 2-nitrophenylphenyl sulfone was dissolved in 100 g of methanol in a 300 ml capacity four-necked flask provided with a stirrer, a thermometer, a dropping funnel and a reflux condensor. An amount of 11.0 g (0.1 mole) of thiophenol and 4.0 g (0.1 mole) of sodium hydroxide were dissolved in 30 g of methanol. The resultant methanol solution of sodium thiophenolate was added to the aforesaid solution of 2-nitrophenylphenyl sulfone, and the reaction was carried out at 60°C for 4 hours.

After the reaction, methanol was concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the organic solvent. The aqueous solution was concentrated to dryness, to provide 16.1 g (97.1 % yield) of sodium benzenesulfinate as white crystals. The purity was found 99.0 % by liquid chromatography.

The removal of toluene from the organic solution by distillation provided 22.6 g (96.7 % yield) of 2-nitrophenylphenyl sulfide as yellow crystals. The purity was found

25

98.8 % by liquid chromatography.

(ii) Oxidation of 2-Nitrophenylphenyl Sulfide and Reuse

An amount of 21.1 g (0.09 mole) of the recovered 2-nitrophenylphenyl sulfide was placed in a 200 ml capacity flask together with 83.3 g of acetic acid and heated to 50°C to provide a uniform solution. The solution was further heated to 85°C, and 26.2 g of a 35 % aqueous solution of hydrogen peroxide (0.27 mole) were dropwise added to the solution of the sulfide over 30 minutes while the solution was maintained at 85°C.

After stirring for 5 hours at 85°C, the resultant reaction mixture was left standing to room temperatures, the resultant precipitates were filtered and dried, to provide 23.0 g (95.4 % yield) of 2-nitrophenylphenyl sulfone, mp. 146.9-148.6°C, which was found 98.2 % in purity by liquid chromatography.

An amount of 21.4 g (0.08 mole) of the thus obtained 2-nitrophenylphenyl sulfone was dissolved in 80 g of methanol, and was reacted with sodium thiophenolate in the same manner as above-mentioned.

After the reaction, methanol was concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the organic solution. The aqueous solution was then concentrated to dryness, to provide 12.7 g (95.8 % yield) of sodium benzenesulfinate as white crystals.

The purity was found 99.1 % by liquid chromatography.

2-Nitrophenylphenyl sulfide was recovered from the toluene solution.

5

EXAMPLE 2

(i) Production of Sodium Benzenesulfinate and Recovery of 2-Nitrophenylphenyl Sulfide

A mixture of 100 g of ethyl acetate and 100 g of water was used as a solvent, and the reaction was carried out in the same manner as in Example 1 in the presence of 0.1 g of benzyltriethylammonium chloride as a phase transfer catalyst.

After the reaction, the reaction mixture was left standing to room temperatures, the resulting aqueous solution was separated from the ethyl acetate solution containing 2-nitrophenylphenyl sulfide, and concentrated to dryness, to provide 13.7 g (82.7 % yield) of sodium benzenesulfinate as white crystals, which was found 99.1 % in purity by liquid chromatography.

The removal of ethyl acetate by distillation from the ethyl acetate solution provided 22.4 g (95.8 % yield) of 2-nitrophenylphenyl sulfide as yellow crystals. The sulfide was found 98.9 % in purity by liquid chromatography.

(ii) Oxidation of 2-Nitrophenylphenyl sulfide and Reuse

An amount of 21.0 g (0.09 mole) of the recovered 2-nitrophenylphenyl sulfide was oxidized with 26.2 g (0.27 mole)

of a 35 % aqueous solution of hydrogen peroxide in the same manner as in Example 1. After the reaction, the reaction mixture was cooled, and resulting precipitates were filtered and dried, to provide 23.1 g (95.9 % yield) of 2-nitrophenyl-
5 phenyl sulfone as white crystals, mp. 147.0-148.7 °C. The purity was found 98.4 % by liquid chromatography.

An amount of 21.4 g (0.08 mole) of the thus obtained 2-nitrophenylphenyl sulfone was dissolved in a mixture of 80 g of ethyl acetate and 80 g of water, and was reacted with
10 sodium thiophenolate in the same manner as in Example 1.

After the reaction, ethyl acetate was concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the organic solution. The aqueous solution was concentrated to dryness, to provide 12.5
15 g (94.2 % yield) of sodium benzenesulfinate as white crystals. The purity was found 99.0 % by liquid chromatography.

2-Nitrophenylphenyl sulfide was recovered from the toluene solution.

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EXAMPLE 3

4-Nitrophenylphenyl sulfone was used in place of 2-nitrophenylphenyl sulfone and otherwise the reaction was carried out in the same manner as in Example 1.

After the reaction, methanol was concentrated, toluene
25 and water were added to the concentrate, and the aqueous

solution was separated from the organic solution containing 4-nitrophenylphenyl sulfide as by-products. The aqueous solution was concentrated to dryness, to provide 14.8 g (88.5 % yield) of sodium benzenesulfinate as white crystals. 5 The purity was found 98.2% by liquid chromatography.

EXAMPLE 4

4-Nitrophenylphenyl sulfone and a solvent as a mixture of 100 g of toluene and 100 g of water were used in place of 10 2-nitrophenylphenyl sulfone and methanol, respectively, and the reaction was carried out in the presence of 0.1 g of tetrabutylammonium hydrogen sulfate as a phase transfer catalyst otherwise in the same manner as in Example 1.

After the reaction, the reaction mixture was left 15 standing to room temperatures, the aqueous solution was separated from the toluene solution containing 4-nitrophenylphenyl sulfide as by-products. The aqueous layer was concentrated to dryness, to provide 14.6 g (88.4 % yield) of sodium benzenesulfinate as white crystals. The purity was found 20 99.4 % by liquid chromatography.

EXAMPLE 5

(i) Production of Sodium Benzenesulfinate and Recovery of 2-Nitrophenylphenyl Sulfide 25 4-Nitrophenylphenyl sulfone and a solvent composed of

a mixture of 100 g of nitrobenzene and 100 g of water were used in place of 2-nitrophenylphenyl sulfone and methanol, respectively, and the reaction was carried out in the presence of 0.1 g of tetrabutylammonium hydrogen sulfate as a phase transfer catalyst otherwise in the same manner as in Example 1.

After the reaction, the reaction mixture was left standing to room temperatures, the aqueous solution was separated from the nitrobenzene solution containing 4-nitrophenylphenyl sulfide. The aqueous layer was concentrated to dryness, to provide 15.3 g (92.5 % yield) of sodium benzene-sulfinate as white crystals. The purity was found 99.3 % by liquid chromatography.

The nitrobenzene solution was concentrated by distillation to dryness, to provide 22.6 g (96.5 % yield) of 4-nitrophenylphenyl sulfide, which was found 98.7 % in purity by liquid chromatography.

(ii) Oxidation of 4-Nitrophenylphenyl sulfide and Reuse

An amount of 21.1 g (0.09 mole) of the recovered 4-nitrophenylphenyl sulfide was oxidized with hydrogen peroxide in the same manner as in Example 1. After the reaction, the reaction mixture was cooled, and resulting precipitates were filtered and dried, to provide 23.1 g (96.0 % yield) of 4-nitrophenylphenyl sulfone as white crystals, mp. 141.4-142.9°C. The purity was found 98.5 % by liquid chromatography.

An amount of 21.4 g (0.08 mole) of the thus obtained 4-nitrophenylphenyl sulfone was dissolved in a two phase solvent composed of a mixture of 80 g of nitrobenzene and 80 g of water, and was reacted with sodium thiophenolate in the same manner as in Example 1.

After the reaction, the resultant aqueous solution was separated from the organic solvent. The aqueous solution was concentrated to dryness, to provide 12.6 g (95.2 % yield) of sodium benzenesulfinate as white crystals. The purity was found 99.2 % by liquid chromatography.

EXAMPLE 6

An amount of 5.6 g (0.1 mole) of potassium hydroxide was used in place of sodium hydroxide, and the reaction was carried out otherwise in the same manner as in Example 1.

After the reaction, methanol was concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the organic solution containing 2-nitrophenylphenyl sulfide. The aqueous solution was concentrated to dryness, to provide 15.0 g (82.4 % yield) of sodium benzenesulfinate as white crystals. The purity was found 99.0 % by liquid chromatography.

EXAMPLE 7

An amount of 30.8 g (0.1 mole) of 2,4-dinitrophenyl-

phenyl sulfone was used in place of 2-nitrophenylphenyl sulfone, and the reaction was carried out otherwise in the same manner as in Example 1.

5 After the reaction, methanol was concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the toluene solution containing 2,4-dinitrophenylphenyl sulfide. The aqueous solution was then concentrated to dryness, to provide 13.8 g (83.4 % yield) of sodium benzenesulfinate as white crystals. The purity was
10 found 99.2 % by liquid chromatography.

EXAMPLE 8

Ethanol was used as a solvent in place of methanol, and the reaction was carried out otherwise in the same manner as
15 in Example 1.

After the reaction, ethanol and methanol were concentrated, toluene and water were added to the concentrate, and the aqueous solution was separated from the toluene solution. The aqueous solution was then concentrated to dryness, to
20 provide 16.0 g (96.7 % yield) of sodium benzenesulfinate as white crystals. The purity was found 99.2 % by liquid chromatography.

The removal of toluene from the toluene solution provided 22.5 g (96.1 % yield) of 2-nitrophenylphenyl sulfide.
25 The purity was found 99.2 % by liquid chromatography.

EXAMPLE 9

An amount of 46.2 g (0.2 mole) of 2-nitrophenylphenyl sulfide was dissolved under heating in 200 g of toluene in a 500 ml capacity four-necked flask provided with a stirrer, a thermometer, a dropping funnel and a reflux condensor, and there were added thereto 1 g of sodium tungstate dihydrate and 2 g of tetrabutylammonium hydrogen sulfate. To the resultant mixture were then dropwise added 40.9 g (0.42 mole) of 30 % by weight aqueous solution of hydrogen peroxide at temperatures of about 80°C under stirring, and the reaction was carried at the temperature for 3 hours. The pH of the reaction mixture was found about 1.

After the completion of the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 51.8 g (97.8 % yield) of 2-nitrophenylphenyl sulfone as pale yellow crystals. The purity was found 99.5 % by liquid chromatography.

EXAMPLE 10

An amount of 46.2 g (0.2 mole) of 2-nitrophenylphenyl sulfide was dissolved under heating in 200 g of toluene in a 500 ml capacity four-necked flask provided with a stirrer, a thermometer, a dropping funnel and a reflux condensor, and

there were added thereto 1 g of sodium tungstate dihydrate, 2 g of benzyltrimethylammonium chloride, and 1.5 g of sulfuric acid in this order. To the resultant mixture were then added dropwise 40.9 g (0.42 mole) of aqueous hydrogen peroxide at temperatures of about 80°C under stirring, and the reaction was carried at the temperature for 3 hours. The pH of the reaction mixture was found about 1.

After the completion of the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 51.6 g (97.6 % yield) of 2-nitrophenylphenyl sulfone as pale yellow crystals. The purity was found 99.5 % by liquid chromatography.

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EXAMPLE 11

Trioctylmethylammonium chloride was used in amounts of 2 g as phase transfer catalysts in place of tetrabutylammonium hydrogen sulfate, and 1.5 g of sulfuric acid were added to a toluene solution of the sulfide together with hydrogen peroxide, and the reaction was carried out otherwise in the same manner as in Example 9.

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 51.8 g (97.4 % yield) of 2-nitrophenylphenyl sulfone as pale

yellow crystals. The purity was found 99.0 % by liquid chromatography.

EXAMPLE 12

5 Benzyldimethylaurylammonium chloride was used in amounts of 2 g as phase transfer catalysts in place of tetrabutylammonium hydrogen sulfate, and 1.5 g of sulfuric acid were added to a toluene solution of the sulfide together with hydrogen peroxide, and the reaction was carried out otherwise
10 in the same manner as in Example 9.

 After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 51.7 g (97.4 % yield) of 2-nitrophenylphenyl sulfone as pale
15 yellow crystals. The purity was found 99.2 % by liquid chromatography.

EXAMPLE 13

 An amount of 46.2 g (0.2 mole) of 4-nitrophenylphenyl
20 sulfide was used in place of 2-nitrophenylphenyl sulfide, and the reaction was carried out otherwise in the same manner as in Example 9.

 After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the
25 toluene solution was concentrated to dryness, to provide

51.9 g (97.7 % yield) of 4-nitrophenylphenyl sulfone as pale yellow crystals. The purity was found 99.1 % by liquid chromatography.

5

EXAMPLE 14

An amount of 46.2 g (0.2 mole) of 4-nitrophenylphenyl sulfide was used in place of 2-nitrophenylphenyl sulfide, and nitrobenzene was used as organic solvents in place of toluene, and the reaction was carried out otherwise in the same manner as in Example 9.

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After the reaction, the reaction mixture was separated into a nitrobenzene solution and an aqueous solution, and the nitrobenzene solution was heated and concentrated, and then cooled. The resultant solids were separated by filtration, washed with methanol, and then dried, to provide 50.3 g (95.2 % yield) of 4-nitrophenylphenyl sulfone as pale yellow crystals. The purity was found 99.7 % by liquid chromatography.

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EXAMPLE 15

An amount of 46.2 g (0.2 mole) of 4-nitrophenylphenyl sulfide was used in place of 2-nitrophenylphenyl sulfide, and 1 g of ammonium molybdate was used in place of sodium tungstate, and the reaction was carried out otherwise in the same manner as in Example 9.

25

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 51.5 g (97.5 % yield) of 4-nitrophenylphenyl sulfone as pale yellow crystals. The purity was found 99.7 % by liquid chromatography.

EXAMPLE 16

An amount of 55.2 g (0.2 mole) of 2,4-dinitrophenylphenyl sulfide was used in place of 2-nitrophenylphenyl sulfide, and the reaction was carried out otherwise in the same manner as in Example 9.

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide 60.4 g (97.3 % yield) of 2,4-dinitrophenylphenyl sulfone as yellow crystals. The purity was found 99.3 % by liquid chromatography.

EXAMPLE 17

An amount of 55.2 g (0.2 mole) of 2,4-dinitrophenylphenyl sulfide was used in place of 2-nitrophenylphenyl sulfide, and 2 g of benzyldimethyltetradecylammonium chloride in place of tetrabutylammonium hydrogen sulfate, and the reaction was carried out otherwise in the same manner as in

Example 9.

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was concentrated to dryness, to provide
5 60.5 g (97.2 % yield) of 2,4-dinitrophenylphenyl sulfone as yellow crystals. The purity was found 99.1 % by liquid chromatography.

EXAMPLE 18

10 Benzyltrimethylammonium chloride was used in amounts of 2 g as phase transfer catalysts in place of tetrabutylammonium hydrogen sulfate, and the reaction was carried out otherwise in the same manner as in Example 9. Since there was added no sulfuric acid to the reaction mixture, the pH
15 of the reaction mixture was found about 7.

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was analyzed by liquid chromatography. The toluene solution was found to contain 2-nitrophenylphenyl
20 sulfone in amounts of 85.6 % and 2-nitrophenylphenyl sulfoxide (by-products) in amounts of 14.4 %.

EXAMPLE 19

25 Benzyltrimethylammonium chloride was used in amounts of 2 g as phase transfer catalysts in place of tetrabutyl-

ammonium hydrogen sulfate, and 1.5 g of sodium hydroxide were added to a toluene solution of the sulfide together with hydrogen peroxide, and the reaction was carried out in an alkaline condition at a pH of about 14 otherwise in the same manner as in Example 9.

After the reaction, the reaction mixture was separated into a toluene solution and an aqueous solution, and the toluene solution was analyzed by liquid chromatography. The toluene solution was found to contain 2-nitrophenylphenyl sulfide (unreacted) in amounts of 73.2 %, 2-nitrophenylphenyl sulfoxide (by-products) in amounts of 15.5 %, and 2-nitrophenylphenyl sulfone in amounts of 11.3 %.

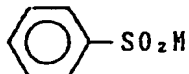
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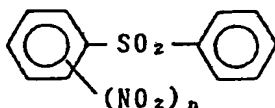
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What Is Claimed Is:

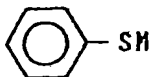
1. A method of producing alkali metal benzenesulfonates having the general formula of



wherein M represents Na or K, which comprises: reacting nitrophenylphenyl sulfones having the general formula of



wherein n is an integer of 1 or 2, with alkali metal thiophenolates having the general formula of



in solvents.

2. The method as claimed in Claim 1 wherein the solvent is a homogeneous one composed of water and water-soluble lower aliphatic alcohols.

3. The method as claimed in Claim 2 wherein the lower aliphatic alcohols are methanol or ethanol.

4. The method as claimed in Claim 1 wherein the reaction is carried out in methanol or ethanol.

5. The method as claimed in Claim 1 wherein the solvent is a two-phase heterogeneous one composed of a mixture of water and water-insoluble or slightly water-soluble organic solvents.

5

6. The method as claimed in Claim 5 wherein the organic solvents are aliphatic hydrocarbons, aromatic hydrocarbons, nitrated aromatic hydrocarbons or carboxylic acid esters.

10

7. The method as claimed in Claim 5 wherein the organic solvents are toluene, nitrobenzene or ethyl acetate.

15

8. The method as claimed in Claim 1 wherein the reaction is carried out in a two-phase heterogeneous solvent composed of a mixture of water and water-insoluble or slightly water-soluble organic solvents in the presence of a phase transfer catalyst.

20

9. The method as claimed in Claim 8 wherein the organic solvents are aliphatic hydrocarbons, aromatic hydrocarbons, nitrated aromatic hydrocarbons or carboxylic acid esters.

10. The method as claimed in Claim 8 wherein the organic solvents are toluene, nitrobenzene or ethyl acetate.

25

11. The method as claimed in Claim 8 wherein the phase transfer catalysts are tetraalkylammoniums or benzyltrialkylammoniums or dibenzyltrialkylammoniums.

5 12. The method as claimed in Claim 8 wherein the phase transfer catalysts are tetraalkylammonium halides or benzyltrialkylammonium halides or dibenzyltrialkylammonium halides.

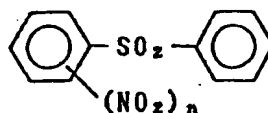
10 13. The method as claimed in Claim 8 wherein the phase transfer catalysts are tetraalkylammonium hydrogen sulfates.

14. The method as claimed in Claim 1 wherein nitrophenylphenyl sulfides having the general formula of



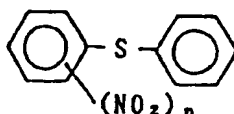
wherein n is an integer of 1 or 2, produced as by-products in the reaction, is oxidized to the nitrophenylphenyl sulfones for reuse in the reaction.

20 15. A method of producing nitrophenylphenyl sulfones having the general formula of



25 wherein n is an integer of 1 or 2, which comprises:

oxidizing nitrophenylphenyl sulfides having the general formula of



5

wherein n is an integer of 1 or 2, with hydrogen peroxide in the presence both of oxidizing catalysts selected from the group consisting of water-soluble tungstates and molybdates and of phase transfer catalysts in a two-phase heterogeneous solvent.

10

16 The method as claimed in Claim 15 wherein the two-phase heterogeneous solvent is composed of water and water-insoluble or slightly water-soluble organic solvents.

15

17. The method as claimed in Claim 15 wherein the tungstates are sodium tungstate, potassium tungstate or ammonium tungstate.

20

18. The method as claimed in Claim 15 wherein the molybdates are sodium molybdate, potassium molybdate or ammonium molybdate.

25

19. The method as claimed in Claim 15 wherein the phase transfer catalysts are tetraalkylammoniums or benzyltrialkylammoniums or dibenzylalkylammoniums.

20. The method as claimed in Claim 15 wherein the phase transfer catalysts are tetraalkylammonium halides or benzyl-trialkylammonium halides or dibenzyltrialkylammonium halides.

5

21. The method as claimed in Claim 15 wherein the phase transfer catalysts are tetraalkylammonium hydrogen sulfates.

22. The method as claimed in Claim 15 wherein the reaction is carried out in acidic or neutral conditions.

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EUROPEAN SEARCH REPORT

0256486

Application number

EP 87 11 1587

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl. 4)
A, D	JOURNAL OF ORGANIC CHEMISTRY, vol. 19, 1954, pages 1704-1707, Easton, US; N. KHARASCH et al.: "Derivatives of sulfenic acids. XV. A new synthesis of thiophenols" * Page 1706, last paragraph *	14, 15	C 07 C 145/00 C 07 C 147/06
			TECHNICAL FIELDS SEARCHED (Int. Cl. 4)
			C 07 C 145/00
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 04-11-1987	Examiner VAN GEYT J. J. A.
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document</p>			